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(54) **Complex oxide having high seebeck coefficient and high electric conductivity**

(57) This invention provides a complex oxide comprising the features of: (i) being represented by the formula:  $\text{Ca}_{3-x}\text{RE}_x\text{Co}_4\text{O}_y$  wherein RE is a rare earth element,  $0 \leq x \leq 0.5$  and  $8.5 \leq y \leq 10$ , (ii) having a Seebeck coefficient of 100  $\mu\text{V/K}$  or more at a temperature

of 300°C or higher, and (iii) having an electric conductivity of  $10^3$  S/m or more at a temperature of 300°C or higher. The complex oxide is composed of low-toxicity elements, excellent in heat resistance and chemical durability and high in thermoelectric conversion efficiency.

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**Description****Field of the Invention**

**[0001]** The present invention relates to a complex oxide having a high Seebeck coefficient and a high electric conductivity, and to a thermoelectric material prepared from said complex oxide.

**Background Art**

**[0002]** In our country, effective energy is obtained at a ratio of only about 30% from the primary supply energy, and about 70% of energy is eventually discarded as heat into the atmosphere. The heat evolved by combustion in plants, garbage-incinerating facilities or the like is thrown away into the atmosphere without conversion into other energy. In this way, we are uselessly casting away a vast amount of thermal energy and are acquiring only a small amount of energy from combustion of fossil fuel or otherwise.

**[0003]** To increase the proportion of acquired energy, the thermal energy to be released into the atmosphere should be effectively utilized. For this purpose, thermoelectric conversion for direct conversion of thermal energy to electrical energy is effective means. The thermoelectric conversion, which utilizes Seebeck effect, is an energy conversion method for generating electricity by creating a difference in temperature between both ends of a thermoelectric material to produce a difference of electric potential. In this thermoelectric generation, electricity is generated simply by setting one end of a thermoelectric material at a location heated to a high temperature by waste heat, and the other end thereof in the atmosphere (room temperature) and connecting a leading wire to both ends thereof. This method entirely eliminates a need for a device with moving parts such as a motor or a turbine required in the common generation of power. As a consequence, the method is economical and can be carried out without giving off a gas by combustion. Moreover, the method can continuously generate electricity until the thermoelectric material is deteriorated.

**[0004]** As set forth above, thermoelectric generation is a technique expected to play a share of role for the resolution of energy problems which will be concerned henceforth. To realize the thermoelectric generation, there is a need for developing thermoelectric materials which have a high thermoelectric conversion efficiency and have excellent heat resistance and chemical durability.

**[0005]** Currently, intermetallic compounds are known as a substance of high thermoelectric conversion efficiency. Among them,  $\text{Bi}_2\text{Te}_3$  has the highest thermoelectric conversion efficiency. However, the thermoelectric conversion efficiency of  $\text{Bi}_2\text{Te}_3$  is only about 10% at highest. Further,  $\text{Bi}_2\text{Te}_3$  can be used only at a temperature of 200°C or lower.  $\text{TeAgSb}$ -containing metallic compound has a high thermoelectric conversion efficiency in a temperature range of about 600 to about 1,000 K which falls in the temperature range of waste heat. However, Te and Sb are rare elements having toxicity and can not be used in the air because they are readily oxidizable. With the above drawbacks,  $\text{Bi}_2\text{Te}_3$  and  $\text{TeAgSb}$ -containing metallic compound are limited in application as a thermoelectric material for use.

**[0006]** In the foregoing situation, it is expected to develop materials composed of low-toxicity elements which have superior heat resistance and chemical durability and a high thermoelectric conversion efficiency.

**[0007]** While metallic oxides may be proposed as materials excellent in heat resistance and chemical durability, the metallic oxides are lower in thermoelectric conversion efficiency by an order of magnitude than  $\text{Bi}_2\text{Te}_3$ . In fact, known oxides having an electric conductivity of  $1 \times 10^3 \text{ S/m}$  or more show a Seebeck coefficient only as low as tens  $\mu\text{V/K}$ .

**Description of the Drawings**

**[0008]** FIG. 1 shows a powder X-ray diffraction pattern of the complex oxide obtained in Example 1. FIG. 2 schematically shows a thermoelectric module produced using the complex oxide of the invention as a thermoelectric material. FIG. 3 is a graph showing the temperature dependency of Seebeck coefficient of the complex oxide prepared in Example 1. FIG. 4 is a graph showing the temperature dependency of electric conductivity of the complex oxide prepared in Example 1. In the drawings, indicated at 1 is a base plate for a high-temperature location; at 2, a base plate for a low-temperature location; at 3, a P-type thermoelectric material; at 4, an N-type thermoelectric material; at 5, an electrode; and at 6, a leading wire.

**Disclosure of the Invention**

**[0009]** A principal object of the present invention is to provide a material composed of low-toxicity elements, the material being excellent in heat resistance and chemical durability and having a high thermoelectric conversion efficiency.

**[0010]** The present inventor conducted extensive research in view of the above-mentioned current situation regarding thermoelectric materials and found that a complex oxide having a specific composition which includes a rare earth

element, Ca, Co and O as constituent elements has a high Seebeck coefficient and a high electric conductivity and is useful as a thermoelectric material in a thermoelectric module. The present invention was completed based on this novel finding.

**[0011]** The present invention provides the following complex oxides and thermoelectric material.

1. A complex oxide comprising the features of:

- (i) being represented by the formula :  $\text{Ca}_{3-x}\text{RE}_x\text{Co}_4\text{O}_y$  wherein RE is a rare earth element,  $0 \leq x \leq 0.5$  and  $8.5 \leq y \leq 10$ ,
- (ii) having a Seebeck coefficient of  $100 \mu\text{V/K}$  or more at a temperature of  $300^\circ\text{C}$  or higher, and
- (iii) having an electric conductivity of  $10^3 \text{ S/m}$  or more at a temperature of  $300^\circ\text{C}$  or higher.

2. The complex oxide as defined in item 1, wherein RE is at least one element selected from the group consisting of Sc, Y, La, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb and Lu.

3. The complex oxide as defined in item 1 which has an oxygen deficient perovskite structure.

4. A P-type thermoelectric material prepared from the complex oxide as defined in item 1.

**[0012]** The complex oxide of the present invention is represented by the formula :  $\text{Ca}_{3-x}\text{RE}_x\text{Co}_4\text{O}_y$ . In the formula, RE is a rare earth element and is at least one element selected from the group consisting of Sc, Y, La, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb and Lu. The value of x ranges from 0.1 to 0.5, preferably from 0.1 to 0.25. The value of y ranges from 8.5 to 10, preferably from 8.8 to 9.2.

**[0013]** Such complex oxide has an oxygen deficient perovskite structure similar to  $\text{Ca}_3\text{Co}_4\text{O}_9$  which is a conventional oxygen deficient perovskite. To clarify this matter, FIG. 1 shows a powder X-ray diffraction pattern of the complex oxide prepared in Example 1 to be described later. In FIG. 1, a pattern (a) shows an X-ray diffraction pattern of the complex oxide of the invention and a pattern (b) indicates an X-ray diffraction pattern of  $\text{Ca}_3\text{Co}_4\text{O}_9$ . From the patterns, it is evident that in the complex oxide of the present invention, the Ca site of  $\text{Ca}_3\text{Co}_4\text{O}_9$  is replaced by a rare earth element in view of absence of a peak resulting from an impurity phase.

**[0014]** The complex oxide having the above-specified composition according to the invention has a Seebeck coefficient of  $100 \mu\text{V/K}$  or more and an electric conductivity of  $10^3 \text{ S/m}$  or more at a temperature of  $300^\circ\text{C}$  or higher. The complex oxide of the invention can exhibit a high thermoelectric conversion efficiency when used as a thermoelectric material for a thermoelectric module, because the oxide has both a high Seebeck coefficient and a high electric conductivity. Further the complex oxide of the invention is excellent in heat resistance and chemical durability and is made of low-toxicity elements so that the oxide is very suitable for use as a thermoelectric material.

**[0015]** The complex oxide of the invention can be prepared by mixing the starting materials in the predetermined proportions and sintering them in an oxidizing atmosphere.

**[0016]** The starting materials are not limited insofar as they can produce the contemplated complex oxide when sintered. Examples are metals, oxides, compounds (such as carbonate), etc. Examples of compounds as the Ca source are calcium oxide ( $\text{CaO}$ ), calcium chloride ( $\text{CaCl}_2$ ), calcium carbonate ( $\text{CaCO}_3$ ), calcium nitrate ( $\text{Ca}(\text{NO}_3)_2$ ), calcium hydroxide ( $\text{Ca}(\text{OH})_2$ ), dimethoxy calcium ( $\text{Ca}(\text{OCH}_3)_2$ ), diethoxy calcium ( $\text{Ca}(\text{OC}_2\text{H}_5)_2$ ), dipropoxy calcium ( $\text{Ca}(\text{OC}_3\text{H}_7)_2$ ), etc. Examples of compounds as the rare earth element (RE) source are oxides ( $\text{RE}_2\text{O}_3$  such as  $\text{Gd}_2\text{O}_3$ ), nitrates ( $\text{RE}(\text{NO}_3)_3$ ), chlorides ( $\text{RECl}_3$ ), hydroxides ( $\text{RE}(\text{OH})_3$ ), alkoxide compounds ( $\text{RE}(\text{OCH}_3)_3$ ,  $\text{RE}(\text{OC}_2\text{H}_5)_3$ ,  $\text{RE}(\text{OC}_3\text{H}_7)_3$ , etc.). Examples of compounds as the Co source are cobalt oxide ( $\text{CoO}$ ,  $\text{Co}_2\text{O}_3$ ,  $\text{Co}_3\text{O}_4$ , etc.), cobalt chloride ( $\text{CoCl}_2$ ), cobalt carbonate ( $\text{CoCO}_3$ ), cobalt nitrate ( $\text{Co}(\text{NO}_3)_2$ ), cobalt hydroxide ( $\text{Co}(\text{OH})_2$ ), dipropoxy cobalt ( $\text{Co}(\text{OC}_3\text{H}_7)_2$ ), etc. Also usable as the starting material are compounds containing at least two elements constituting the complex oxide of the invention.

**[0017]** Sintering means are not limited and can be selected from any desired means such as electric furnaces, gas furnaces and the like, and sintering may be conducted in an oxygen stream, the air or like oxidizing atmosphere.

**[0018]** The sintering temperature and sintering time are not limited and are so determined as to meet the conditions which allow the production of contemplated complex oxide. For example, the sintering is performed at about  $920$  to about  $1100^\circ\text{C}$  for about 20 to about 40 hours. When carbonate, an organic compound or the like is used as the starting material, preferably the starting materials are calcined for decomposition prior to sintering and then sintered to give the contemplated complex oxide. For example, when carbonate is used as the starting material, it is calcined at about  $800$  to about  $900^\circ\text{C}$  for about 10 hours and then sintered under the above-mentioned conditions.

**[0019]** The amount of oxygen in a complex oxide to be produced is controllable depending on the partial pressure of oxygen in sintering, sintering temperature, sintering time and the like. The higher the partial pressure of oxygen is, the higher the value of y in said formula is.

**[0020]** The complex oxide of the invention prepared in this way has both a high Seebeck coefficient and a high electric conductivity so that the oxide can be effectively used as the thermoelectric material for a thermoelectric module.

[0021] FIG.2 schematically shows by way of example a thermoelectric module produced using the complex oxide of the invention as the thermoelectric material. Said thermoelectric module has the same structure as conventional thermoelectric modules. The complex oxide of the invention is used as a P-type thermoelectric material in a thermoelectric module which comprises a base plate for a high-temperature location 1, a base plate for a low-temperature location 2, a P-type thermoelectric material 3, an N-type thermoelectric material 4, an electrode 5 and a leading wire 6.

[0022] The complex oxide of the invention has a high Seebeck coefficient and a high electric conductivity and is excellent in heat resistance and chemical durability.

[0023] The complex oxide of the invention finds applications as a thermoelectric material usable at a high temperature unlike conventional intermetallic compounds which are unsuitable for such applications. Consequently it is expected that the thermal energy heretofore cast away into the atmosphere can be effectively used by incorporating the complex oxide of the invention into a thermoelectric generation system.

#### Examples

[0024] Examples are given below to further clarify the features of the present invention.

#### Example 1

[0025] Using calcium carbonate ( $\text{CaCO}_3$ ) as the Ca source, gadolinium oxide ( $\text{Gd}_2\text{O}_3$ ) as the Gd source, and cobalt oxide ( $\text{Co}_3\text{O}_4$ ) as the Co source, thorough mixing was conducted at a mixing ratio (mole ratio) of 2.5 : 9.5 : 4 of Ca : Gd : Co. The mixture was charged into a crucible made of alumina and was calcined in an electric furnace at a temperature of 800°C for 10 hours. The calcined product was crushed, press-molded and sintered in an oxygen stream at 920°C for 40 hours to produce a complex oxide. The obtained complex oxide is one represented by the formula  $\text{Ca}_{2.5}\text{Gd}_{0.5}\text{Co}_4\text{O}_9$ .

[0026] FIG.3 is a graph showing the temperature dependency of Seebeck coefficient (S) of the obtained complex oxide at 100 to 700°C. In FIG.3, the mark ● represents the measured results of the complex oxide of the invention and the mark ■ represents the measured results of  $\text{Ca}_3\text{Co}_4\text{O}_9$ . From FIG.3, it is clear that the obtained complex oxide exhibits a Seebeck coefficient of 100  $\mu\text{V/K}$  or more in the temperature range of 300 to 700°C.

[0027] FIG.4 is a graph showing the temperature dependency of electric conductivity ( $\sigma$ ) of said complex oxide as determined by a direct current 4-terminal method. It is apparent from FIG.4 that the electric conductivity of the complex oxide increases with the elevation of temperature, i.e., shows a semiconductor-like behavior and reaches a high value of more than  $3 \times 10^3 \text{ S/m}$  at 300°C. Examples 2-16

[0028] Complex oxides represented by the formula  $\text{Ca}_{2.5}\text{RE}_{0.5}\text{Co}_4\text{O}_{8.8-9.2}$  were produced in the same manner as in Example 1 with the exception of replacing the rare metal element compound used as the starting material with the compounds shown in Table 1.

[0029] In the examples, the Seebeck coefficient was calculated from a graph showing the relationship between thermoelectric voltage and temperature difference as measured using a laboratory-designed instrument and the electric conductivity was measured using a standard four-probe method. The Seebeck coefficient and the electric conductivity were measured in air.

[0030] The Seebeck coefficient (S) and electric conductivity ( $\sigma$ ) of the obtained complex oxides were measured with the results shown in Table 1.

Table 1

Ex. No.	RE		Seebeck Coefficient ( $\mu\text{V/K}$ )	Electric Conductivity ( $10^3\text{S/m}$ )	Measuring Temperature ( $^\circ\text{C}$ )
	Kind	Starting Compound			
1	Gd	Gd <sub>2</sub> O <sub>3</sub>	170	4.3	600
2	Sc	Sc <sub>2</sub> O <sub>3</sub>	152	4.8	700
3	Y	Y <sub>2</sub> O <sub>3</sub>	163	4.5	500
4	La	La <sub>2</sub> O <sub>3</sub>	145	4.2	500
5	Ce	CeO <sub>2</sub>	140	3.8	500
6	Pr	Pr <sub>2</sub> O <sub>3</sub>	155	4.5	500
7	Nd	Nd <sub>2</sub> O <sub>3</sub>	160	4.6	500
8	Sm	Sm <sub>2</sub> O <sub>3</sub>	154	4.0	500
9	Eu	Eu <sub>2</sub> O <sub>3</sub>	120	5.2	600
10	Tb	Tb <sub>2</sub> O <sub>3</sub>	135	3.6	600
11	DY	DY <sub>2</sub> O <sub>3</sub>	147	3.9	600
12	Ho	Ho <sub>2</sub> O <sub>3</sub>	150	4.1	700
13	Er	Er <sub>2</sub> O <sub>3</sub>	158	4.2	700
14	Tm	Tm <sub>2</sub> O <sub>3</sub>	142	4.5	700
15	Yb	Yb <sub>2</sub> O <sub>3</sub>	139	4.8	700
16	Lu	Lu <sub>2</sub> O <sub>3</sub>	120	5.3	700

Examples 17-32

[0031] Complex oxides represented by the formula  $\text{Ca}_{2.75}\text{RE}_{0.25}\text{Co}_4\text{O}_{8.8-9.2}$  were produced in the same manner as

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in Example 1 with the exception of replacing the rare metal element compound used as the starting material with the compounds shown in Table 2, using the starting materials in a mixing mole ratio of 2.75 : 0.25 : 4 of Ca: RE : Co and sintering them at 1000°C for 30 hours.

[0032] The Seebeck coefficient (S) and electric conductivity ( $\sigma$ ) of the obtained complex oxides were measured with the results shown in Table 2.

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Table 2

Ex. No.	RE		Seebeck Coefficient (μV/K)	Electric Conductivity (10 <sup>3</sup> S/m)	Measuring Temperature (C°)
	Kind	Starting Compound			
17	Gd	Gd <sub>2</sub> O <sub>3</sub>	157	5.2	600
18	Sc	Sc <sub>2</sub> O <sub>3</sub>	146	5.6	700
19	Y	Y <sub>2</sub> O <sub>3</sub>	160	5.3	500
20	La	La <sub>2</sub> O <sub>3</sub>	140	5.8	500
21	Ce	CeO <sub>2</sub>	138	5.8	500
22	Pr	Pr <sub>2</sub> O <sub>3</sub>	142	5.5	500
23	Nd	Nd <sub>2</sub> O <sub>3</sub>	156	5.4	500
24	Sm	Sm <sub>2</sub> O <sub>3</sub>	150	5.5	500
25	Eu	Eu <sub>2</sub> O <sub>3</sub>	115	6.0	600
26	Tb	Tb <sub>4</sub> O <sub>7</sub>	127	5.8	600
27	Dy	Dy <sub>2</sub> O <sub>3</sub>	136	5.7	600
28	Ho	Ho <sub>2</sub> O <sub>3</sub>	145	5.5	700
29	Er	Er <sub>2</sub> O <sub>3</sub>	145	5.4	700
30	Tm	Tm <sub>2</sub> O <sub>3</sub>	130	5.9	700
31	Yb	Yb <sub>2</sub> O <sub>3</sub>	132	5.9	700
32	Lu	Lu <sub>2</sub> O <sub>3</sub>	115	6.3	700

## Examples 33-48

5 [0033] Complex oxides represented by the formula  $\text{Ca}_{2.9}\text{RE}_{0.1}\text{Co}_4\text{O}_{8.9-9.2}$  were produced in the same manner as in Example 1 with the exception of replacing the rare metal element compound used as the starting material with the compounds shown in Table 3, using the starting materials in a mixing mole ratio of 2.9 : 0.1 : 4 of Ca: RE : Co and sintering them at 1100°C for 20 hours.

[0034] The Seebeck coefficient (S) and electric conductivity ( $\sigma$ ) of the obtained complex oxides were measured with the results shown in Table 3.

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Table 3

Ex. No.	RE		Seebeck Coefficient ( $\mu\text{V/K}$ )	Electric Conductivity ( $10^3\text{S/m}$ )	Measuring Temperature ( $^\circ\text{C}$ )
	Kind	Starting Compound			
33	Gd	$\text{Gd}_2\text{O}_3$	150	6.5	600
34	Sc	$\text{Sc}_2\text{O}_3$	140	6.8	700
35	Y	$\text{Y}_2\text{O}_3$	155	6.6	500
36	La	$\text{La}_2\text{O}_3$	128	7.2	500
37	Ce	$\text{CeO}_2$	128	7.3	500
38	Pr	$\text{Pr}_2\text{O}_3$	130	7.0	500
39	Nd	$\text{Nd}_2\text{O}_3$	147	6.7	500
40	Sm	$\text{Sm}_2\text{O}_3$	137	7.0	500
41	Eu	$\text{Eu}_2\text{O}_3$	105	7.8	600
42	Tb	$\text{Tb}_2\text{O}_3$	114	7.7	600
43	Dy	$\text{Dy}_2\text{O}_3$	127	7.1	600
44	Ho	$\text{Ho}_2\text{O}_3$	136	6.9	700
45	Er	$\text{Er}_2\text{O}_3$	137	6.7	700
46	Tm	$\text{Tm}_2\text{O}_3$	123	7.0	700
47	Yb	$\text{Yb}_2\text{O}_3$	120	7.1	700
48	Lu	$\text{Lu}_2\text{O}_3$	102	8.0	700

[0035] The results obtained above show that the complex oxides of the present invention have a high Seebeck coefficient and a high electric conductivity.

## Claims

1. A complex oxide comprising the features of:

- 5 (i) being represented by the formula :  $\text{Ca}_{3-x}\text{RE}_x\text{Co}_4\text{O}_y$  wherein RE is a rare earth element,  $0 \leq x \leq 0.5$  and  $8.5 \leq y \leq 10$ ,  
(ii) having a Seebeck coefficient of  $100 \mu\text{V/K}$  or more at a temperature of  $300^\circ\text{C}$  or higher, and  
(iii) having an electric conductivity of  $10^3 \text{ S/m}$  or more at a temperature of  $300^\circ\text{C}$  or higher.
- 10 2. The complex oxide according to claim 1, wherein RE is at least one element selected from the group consisting of Sc, Y, La, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb and Lu.
3. The complex oxide according to any one of claims 1 or 2 which has an oxygen deficient perovskite structure.
- 15 4. A P-type thermoelectric material prepared from the complex oxide as defined in any one of claims 1 to 3.

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FIG. 1

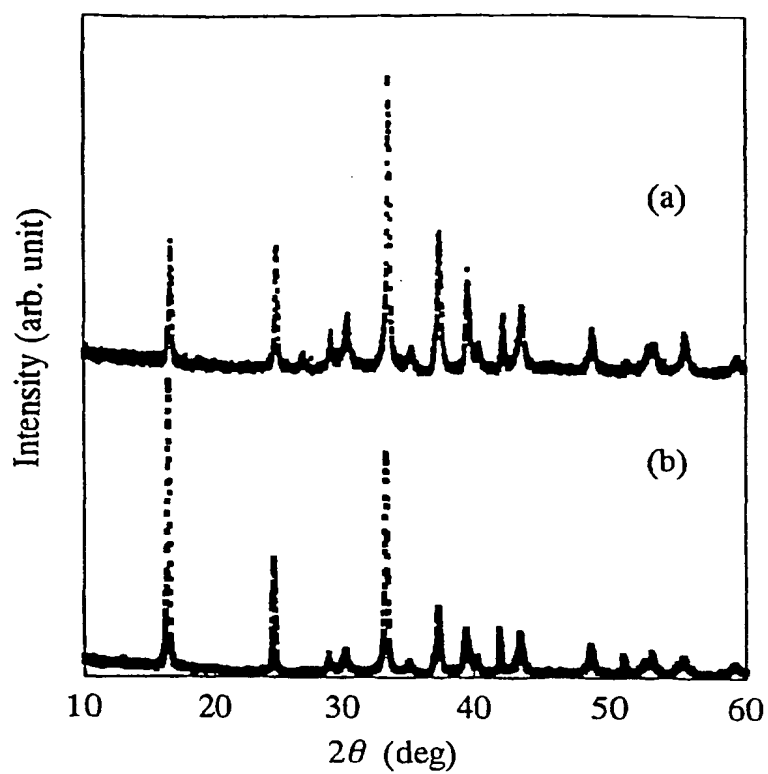


FIG. 2

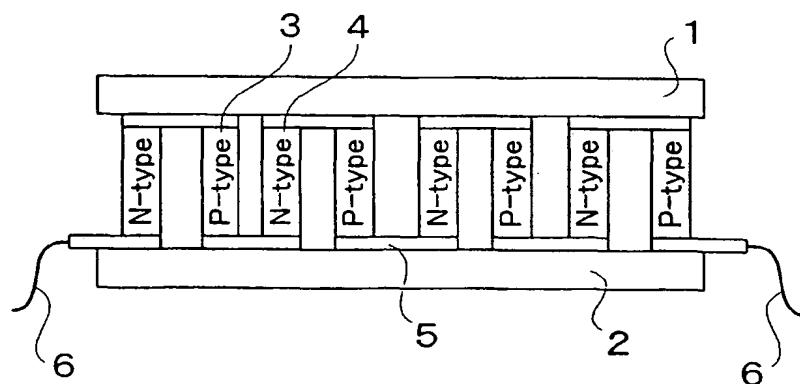
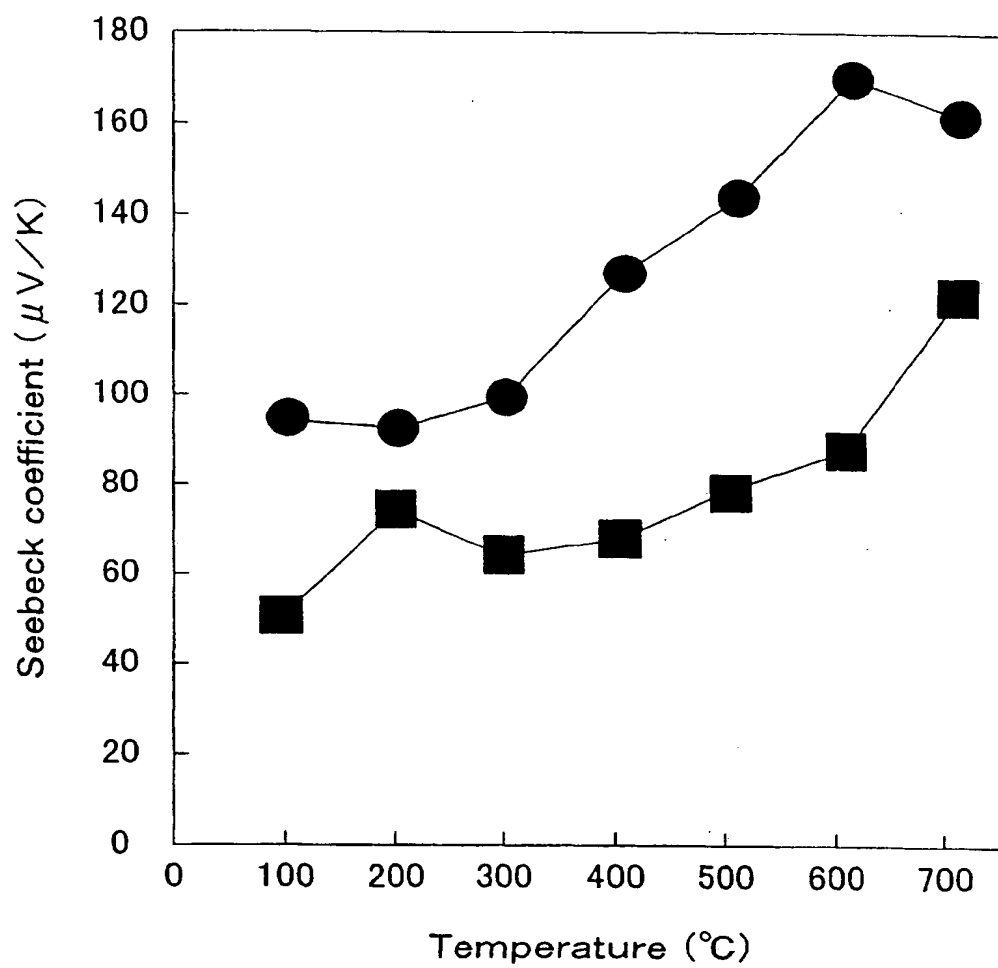
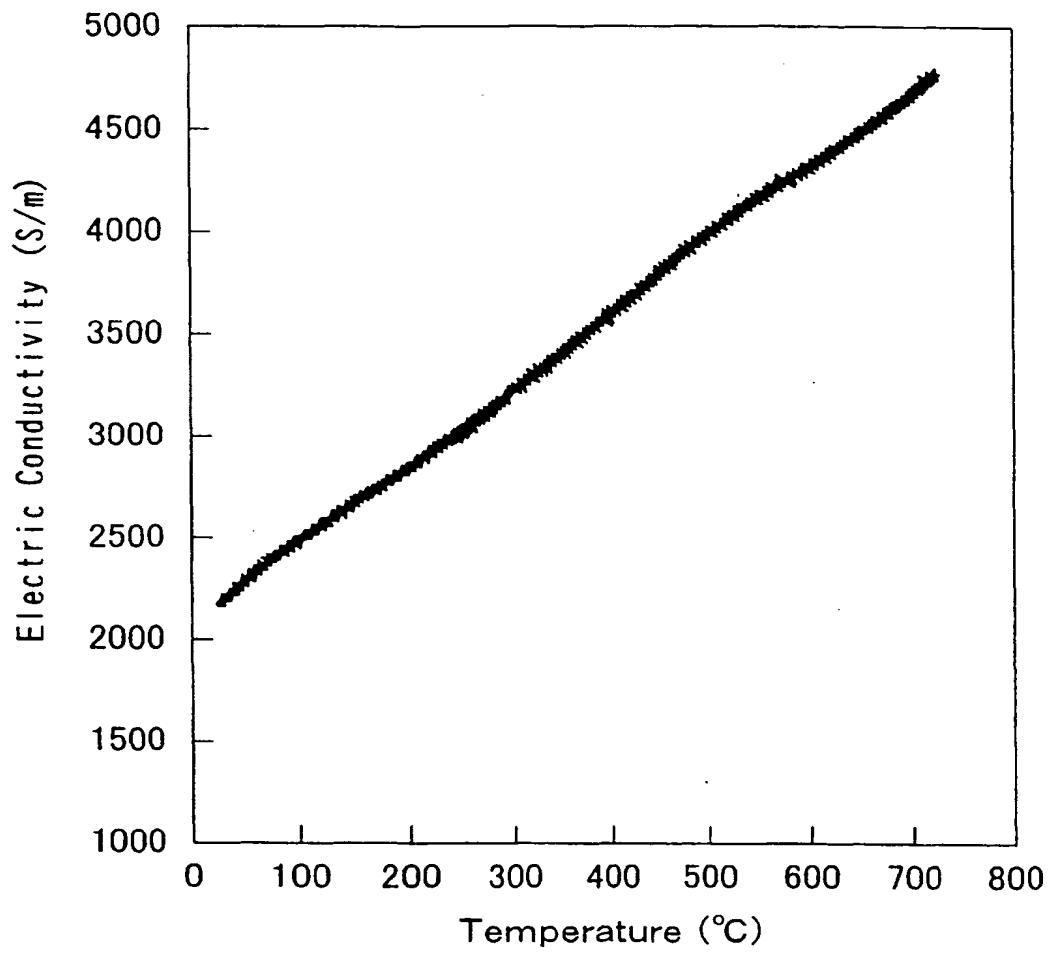


FIG. 3



*FIG. 4*



European Patent  
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# EUROPEAN SEARCH REPORT

Application Number  
EP 00 12 6568

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
X	EP 0 602 782 A (SANYO ELECTRIC CO) 22 June 1994 (1994-06-22) * page 2 *	1	C01G51/00 H01L35/14 H01B1/08
X	US 5 168 095 A (MUNAKATA FUMIO ET AL) 1 December 1992 (1992-12-01) * claims 1-27 *	1,2	
X	DATABASE CHEMABS 'Online! CHEMICAL ABSTRACTS SERVICE, COLUMBUS, OHIO, US; LI, S. ET AL: "High temperature thermoelectric properties of oxide Ca9Co12O28" retrieved from STN Database accession no. 134:201219 CA XP002168496 * abstract * & PROC. INT. CONF. THERMOELECTR. (1999), 18TH, 581-583 , 1999,	1,3	
X,P	DATABASE WPI Section Ch, Week 200043 Derwent Publications Ltd., London, GB; Class L03, AN 2000-485045 XP002168497 & JP 03 069701 B (AGENCY OF IND SCI & TECHNOLOGY), 24 July 2000 (2000-07-24) * abstract *	1,3,4	C01G H01L H01B
The present search report has been drawn up for all claims			
Place of search <b>THE HAGUE</b>		Date of completion of the search <b>6 June 2001</b>	Examiner <b>LIBBERECHT, E</b>
<p>CATEGORY OF CITED DOCUMENTS</p> <p>X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document</p> <p>T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons A : member of the same patent family, corresponding document</p>			

EPC-FCM 1503 03/02 (P-3400)

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ON EUROPEAN PATENT APPLICATION NO.**

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This annex lists the patent family members relating to the patent documents cited in the above-mentioned European search report. The members are as contained in the European Patent Office EDP file on  
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06-06-2001

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
EP 0602782      A	22-06-1994	JP 6163080 A	10-06-1994
		DE 69310689 D	19-06-1997
		DE 69310689 T	18-12-1997
		US 5601949 A	11-02-1997
		US 5670275 A	23-09-1997
US 5168095      A	01-12-1992	JP 2066268 C	24-06-1996
		JP 2074505 A	14-03-1990
		JP 7096443 B	18-10-1995
JP 3069701      B	17-11-1989	JP 1286823 A	17-11-1989
		JP 1705027 C	27-10-1992

EPC FORM PUA/88

For more details about this annex : see Official Journal of the European Patent Office, No. 12/82